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Thermal Treatment of Polyvinyl Alcohol for Coupling MoS₂ and TiO₂ Nanotube Arrays toward Enhancing Photoelectrochemical Water Splitting Performance

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Abstract: Solar-driven photoelectrochemical (PEC) water splitting, using semiconductor photoelectrodes, is considered a promising renewable energy source and solution for environmental sustainability. Herein, we report polyvinyl alcohol (PVA) as a binder material for combining MoS₂ and TiO₂ nanotube arrays (TNAs) to improve PEC water splitting ability. By a thermal treatment process, the formation of the π junction in the PVA structure enhanced the PEC performance of MoS₂/TNAs, exhibiting linear sweeps in an anodic direction with the current density over 65 $\mu\text{A}/\text{cm}^2$ at 0 V vs. Ag/AgCl. Besides, the photoresponse ability of MoS₂/TNAs is approximately 6-fold more significant than that of individual TNAs. Moreover, a Tafel slope of 140.6 mV/decade has been obtained for the oxygen evolution reaction (OER) of MoS₂/TNAs materials.

Keywords: photoelectrochemical; polyvinyl alcohol; oxygen evolution reaction; MoS₂/TiO₂ nanotube arrays



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1. Introduction

Photoelectrochemical (PEC) water splitting based on semiconductor photoanodes has rapidly developed as an alternative energy production approach due to its simple fabrication process and environmental friendliness [1–4]. Fujishima and Honda first discovered the water-splitting in 1972 over TiO₂ nanomaterials for the hydrogen evolution reaction (HER) and oxygen evolution reaction (OER). Since then, many studies have focused on developing and improving semiconductor photoelectrode performances and increasing large-scale synthesizing for PEC water splitting [5]. TiO₂ nanotube arrays (TNAs) are often modified to extend the photoresponse range by combining with a noble metal, creating self-defects, and coupling with other semiconductors to improve the material working efficiency [6,7]. Especially, TNAs modified by MoS₂ demonstrated superior performance in the water-splitting process due to their suitable redox for both HER and OER [8–10]. There are many routes to modify TNAs by MoS₂ that demonstrated promising results, i.e., the hydrothermal method [11,12], physical vapor deposition [13], and chemical vapor deposition [14]. However, they are challenging for industrial-scale applications because of their complex synthesis conditions and small synthesis scale.

Polyvinyl alcohol (PVA) is an inexpensive and widely used polymer. It is readily synthesized in the form of a semicrystalline polymer, which has remarkable properties such as good film-forming ability, thermal and mechanical stability, etc. [15]. In addition, the substantial dispersion of PVA allows it to perform as a proton exchange membrane in fuel cell applications [16,17]. However, PVA is a nonconductive polymer, making it not

easily applied for photoelectronic fields. Modification of PVA to form a conductive polymer throughout π conjugated bonds was proposed as a potential solution. The conjugated PVA (cPVA) connects p orbitals with delocalized electrons in a molecule, lowering the overall energy of the molecule and increasing its stability. In addition, the conjugation is formed by the interaction of unhybridized p atomic orbitals on atoms employing sp^2 - and sp -hybridization, ensuring that the π system of the molecule is created [18,19]. Therefore, some physical conditions, such as annealing temperature treatment or chemical modifications with the support of cross-linkers, may be employed to generate cPVA from commercial PVA [20–22].

In this study, we developed a new strategy for combining MoS_2 and TNAs by using the cPVA binding agent that is formed by a low thermal treatment process. To the best of our knowledge, there is no report on using cPVA to combine two semiconductor materials to enhance PEC performance. The formation of cPVA and its effect on the PEC ability of MoS_2 /TNAs is also clearly explained.

2. Results and Discussion

The FTIR technique was used to evaluate the formation of the π conjugated network in the PVA structure with the C=O and C=C vibration signals. Figure 1 reveals the FTIR spectra of pure PVA and cPVA treated by thermal route from 100 °C to 180 °C. In general, the FTIR spectra of all samples show specific peaks centered at 3400 cm^{-1} and 1725 cm^{-1} corresponding to the stretching vibration of the OH group and the C=O stretching, respectively [23]. In addition, the peaks in the 940–670 cm^{-1} region correspond to the C=C vibration, while the peaks at 1320–1030 cm^{-1} region correspond to C–O stretching vibrations [23]. The peaks located at the region from 2944 cm^{-1} to 2907 cm^{-1} are characterized for stretching vibration of C–H on the backbone of PVA [23,24]. As can be observed, the backbone vibration of PVA remains even after thermal treatment from 100 °C to 160 °C. The peak intensity of the C=O vibration is significantly enhanced, whereas the peak intensity of the O–H vibration is reduced by increasing the thermal treatment temperature from 100 °C to 160 °C. Furthermore, the vibration around 1435 cm^{-1} corresponds to the bending vibration of the OH radical that directly bonds to the C=O radicals in the cPVA structure, which distinguishes it from intramolecular vibration [22,23]. These results indicated that the formation of acetal bridge (C=O bonds) is directly related to the decrease in OH by thermal treatment. However, vibrations such as C–OH, and C=O are very low or not obtained in pristine PVA. The intensity of C=O vibrations increases with the increasing temperature of the thermal treatment and reaches the highest intensity at 140 °C following the decrease in OH intensity.

The absorbance ratio of C=O and OH vibration was calculated for the thermally treated PVA from 100 °C to 160 °C to support the abovementioned argument. Thus, the peak intensity ratio of C=O and OH vibration for the PVA@100 is about 0.71. Meanwhile, PVA@120 is about 0.81, which is much higher than PVA@140 (1.28). However, when the treated temperature is increased up to 160 °C, the peak intensity of OH vibration is decreased without increasing C=O vibration intensity. Specifically, the peak intensity ratio of the C=O vibration and OH vibration of PVA@160 is around 0.88, much lower than PVA@140. When the temperature reaches 180 °C, the structure of PVA seems to be decomposed. The vibration of the OH groups in a range of 3700–3000 cm^{-1} cannot be observed clearly, and the vibration of C=O, HO–C bonds almost disappears. The remained vibrations are observed at 842 cm^{-1} , corresponding to the C=C bond, and a band in the range of 1026–1430 cm^{-1} is attributed to the C–O vibration with a very weak signal. Besides, the colorless state of PVA had been changed into deep yellow after undergoing thermal treatment at 180 °C. All these phenomena above support the hypothesis that self-conjunction occurs through the dehydration process of the OH groups on the PVA chain to form the π conjunction network.

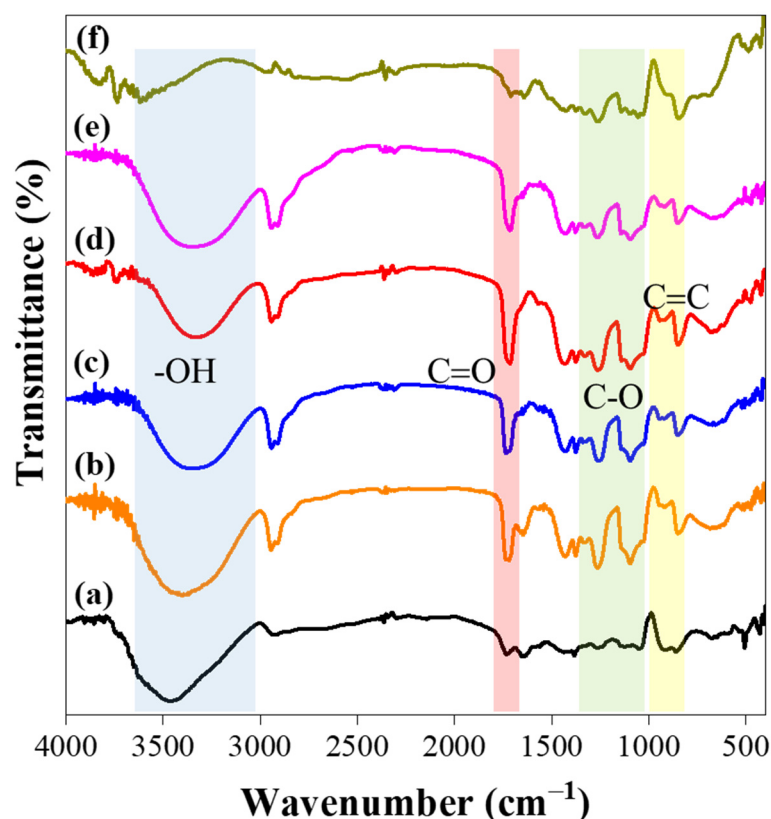


Figure 1. FTIR spectra of PVA (a) and cPVA at 100 °C (b), 120 °C (c), 140 °C (d), 160 °C (e), and 180 °C (f).

Figure 2 shows the cyclic voltammograms (CV) curves under dark condition of samples that were polarized from the resting potential to 1 V and back to -1 V vs. Ag/AgCl in 1 M Na_2SO_4 electrolyte solution. As in previous publications, the specific surface area of TNAs is high due to their tubular structures that contribute on an extension of spacing for the rapid diffusion of electrolyte ions leading to an increase the capacitive performance [25,26]. As a result, a characteristic anodic peak in the range $-1 < E < -0.5$ V appeared relating to the reduction of Ti^{4+} to Ti^{3+} ions, indicating that the high density of Ti^{3+} states causes low resistances and that an increase in current occurred [27,28]. In addition, a cathodic range corresponds to the H^+ depolarization region of these materials, expressing more electrochemical activity that was observed [29,30]. Obviously, the electrochemical activity of $\text{MoS}_2@\text{PVA}/\text{TNAs}@140$ is the highest and the anodic peak of this sample is shifted from -0.7 V vs. Ag/AgCl to -0.5 V vs. Ag/AgCl compared to TNAs. The outstanding electrochemical activity of $\text{MoS}_2@\text{PVA}/\text{TNAs}@140$ composites is also shown through the difference in their CV characteristics with the individual materials such as TNAs, MoS_2 , and cPVA (Figure S1). Moreover, FTIR results (Figure 1) meant that it could be concluded that a high temperature could destroy the structure of PVA, the π conjugation will be destructed, leading to the vanishment of the current density [31].

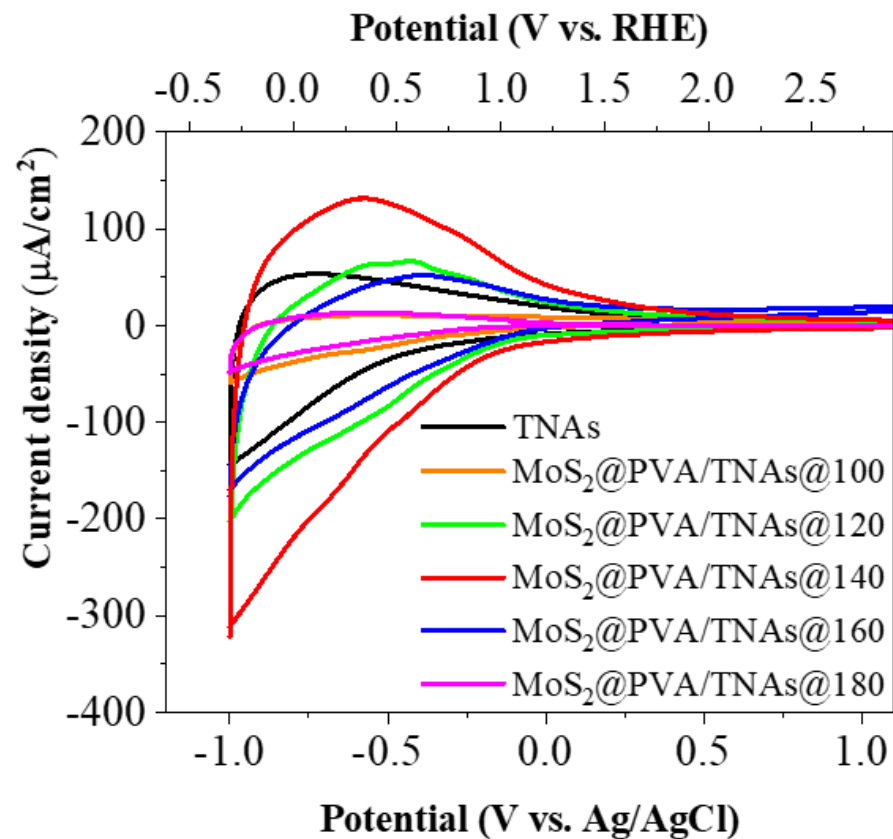


Figure 2. CV curves of MoS₂@PVA/TNAs at different thermal treatment conditions of PVA. (Scan rate: 10 mV s⁻¹).

The XRD patterns of materials are measured and shown in Figure 3. The diffraction peaks of the pristine TNAs (black line) can be easily observed at $2\theta = 25.43^\circ$, 38.08° , 54.15° , and 55.17° corresponding to the planes of (101), (004), (105), and (211) of the anatase phase, respectively (JCPDS No. 88–1175). The XRD pattern of MoS₂ (purple line) shows the crystallization and single phase of MoS₂ with the hexagonal crystal structure. In detail, the diffraction peaks (2θ) positioned at 13.96° and 33.49° , correspond to the (002) and (100) planes of the 2H–MoS₂ phase, respectively (JCPDS No. 37–1492). The XRD pattern of the pure PVA (green line) exhibits characteristic peaks at 19.55° and 22.70° , corresponding to the (101) and (210) planes, respectively. Besides, a weak diffraction peak is also observed at $2\theta = 40.5^\circ$, which is related to the semi-crystalline nature of PVA [32]. The XRD result of the MoS₂@PVA/TNAs@140 sample (red line) shows that three diffraction peaks at $2\theta = 19.85^\circ$, 25.45° , and 33.5° , corresponding to cPVA, (100) plane of 2H–MoS₂ phase, and (101) plane of anatase TiO₂, respectively. The diffraction peak position at $2\theta = 19.85^\circ$ observed at MoS₂@PVA/TNAs@140 diffractogram is supposed to correspond to the formation of C=O or C=C bonds in the cPVA [33]. This result confirmed the formation of cPVA after being treated at 140 °C.

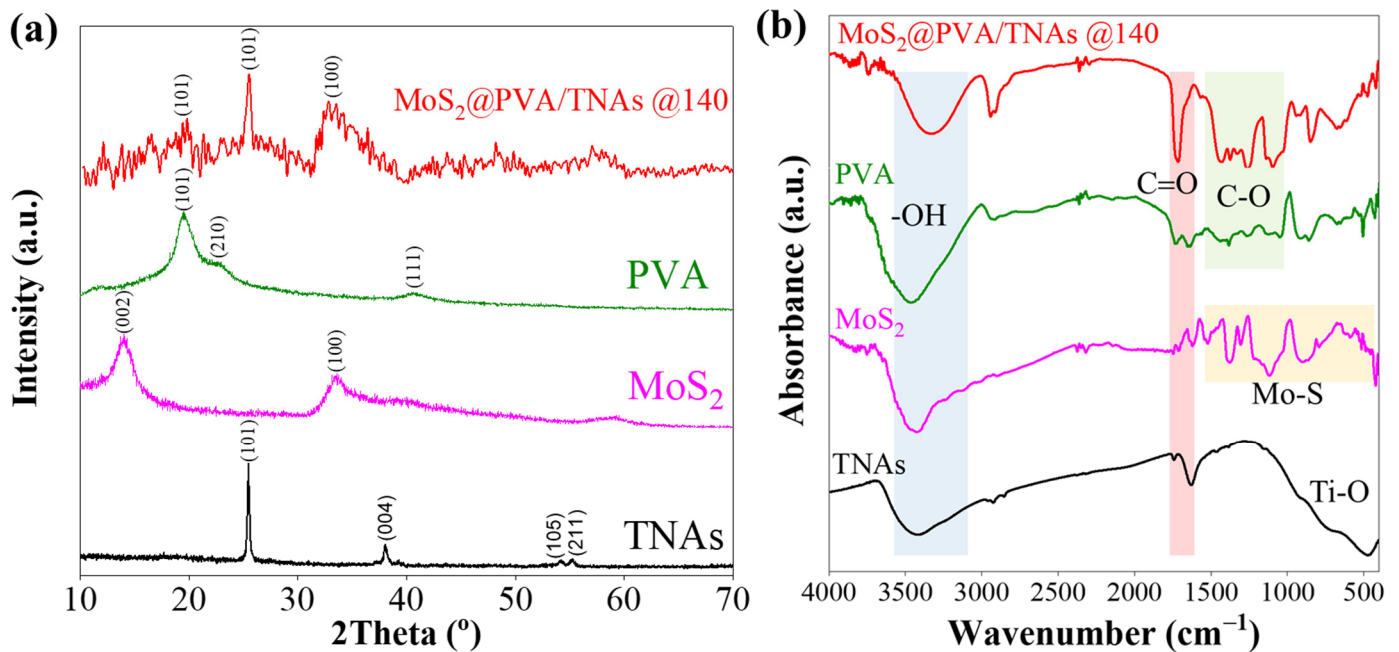


Figure 3. XRD patterns (a) and FTIR spectra (b) of TNAs, MoS₂, PVA, and MoS₂@PVA/TNAs@140.

Furthermore, the FTIR spectra of TNAs, MoS₂, PVA, and MoS₂@PVA/TNAs@140 are measured to confirm the vibration of bonds and are shown in Figure 3b. A typical vibration peak at 475 cm⁻¹ represents the Ti–O–bond in TNAs, and characteristic vibrations throughout many broad absorption bands from 1623 cm⁻¹ to 430 cm⁻¹ are seen of the MoS₂ sample, attributed to the Mo–S vibration [34,35]. The FTIR spectrum of pristine PVA shows strong absorption bands at the wavenumbers of 2930 cm⁻¹, 1725 cm⁻¹, 1640 cm⁻¹, 1440, and 1030 cm⁻¹ to 1320 cm⁻¹ range, corresponding to the stretching vibration of the C–H on the backbone of PVA, C=O stretching, C=C stretching, C–H bending, and C–O stretching, respectively [36]. The formation of conjugated bonds in the cPVA structure was confirmed by a decrease in the –OH absorbance intensity and a significant increase in the intensity of the characteristic peak of the C=O group at 1725 cm⁻¹ in MoS₂@PVA/TNAs@140 sample. This result is consistent with the result for the effect of thermal treatment on PVA seen in Figure 1.

Figure 4 shows SEM images of materials depicting the morphology and microstructure of the TNAs, MoS₂, and MoS₂@PVA/TNAs@140 samples. The results demonstrate that the average diameter of the tubular TiO₂ is about 80–100 nm (Figure 4a), while the MoS₂ has many layers stacked on top of each other, similar to fish-flake structures (Figure 4b). As shown in Figure 4c, TNAs were covered by additional material layers attributed to the MoS₂ in the cPVA matrix. In addition, the difference of the MoS₂ morphology could be explained due to the contribution of cPVA after the thermal treatment process, as seen in Figure 4b,c.

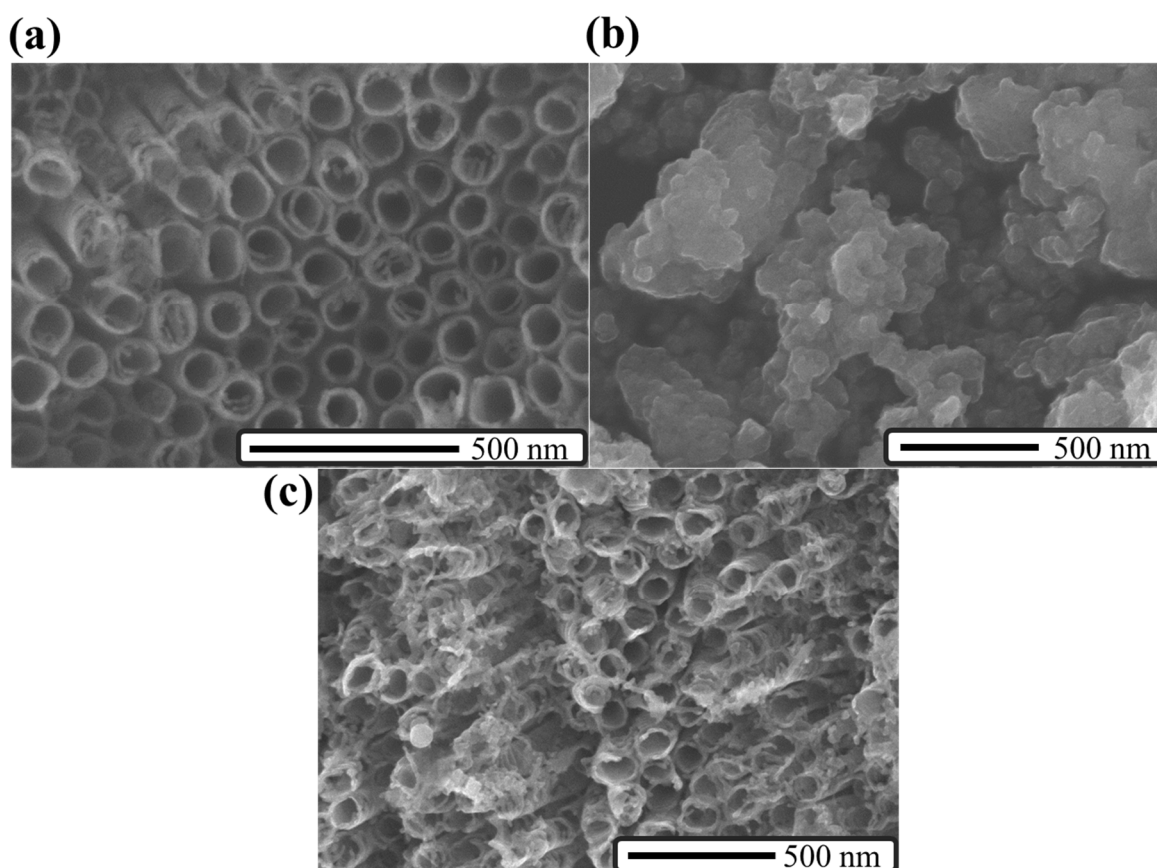


Figure 4. SEM images of TNAs (a), MoS₂ (b), and MoS₂@PVA/TNAs@140 (c).

The LSV of the materials is measured to evaluate the PEC performance. The photoresponse ability of the materials was then evaluated with an applied potential at 0 V vs. Ag/AgCl. Figure 5a shows that the current density of TNAs versus applied potential is higher under visible light than that of dark conditions. In detail, the onset potential of TNAs is obtained at -0.42 V vs. Ag/AgCl in the dark, while it is achieved at -0.55 V vs. Ag/AgCl with the highest current density ($11.2 \mu\text{A}$) under visible light. Besides, the LSV results of MoS₂ in both dark and light conditions are not significantly different. The onset potential was obtained at -0.2 V vs. Ag/AgCl. On the other hand, the combination of MoS₂ and TNAs with a cPVA binding agent indicated an efficient result for OER ability compared to other materials such as TNAs, MoS₂, and cPVA (Figure S2). In detail, the onset potential of the MoS₂@PVA/TNAs@140 sample is obtained at around -0.78 V vs. Ag/AgCl. A current density slope increases from -0.54 V vs. Ag/AgCl and readily reaches saturation at around 0.25 V vs. Ag/AgCl with a current density of about $83.4 \mu\text{A}/\text{cm}^2$. Furthermore, Figure 5b shows that the photoresponse ability of MoS₂ is the lowest, with a different current density between dark and light conditions of just about $1.0 \mu\text{A}/\text{cm}^2$. In contrast, the current density of TNAs is around $11.0 \mu\text{A}/\text{cm}^2$, but it is not stable after five cycles. Specifically, the current density of MoS₂@PVA/TNAs@140 surpasses $65.0 \mu\text{A}/\text{cm}^2$ at 0 V vs. Ag/AgCl in the first cycle of visible light irradiation, which is double that of MoS₂/TNAs@140. Moreover, even after five cycles of “on-off” visible light irradiation, the photoresponse ability of the MoS₂@PVA/TNAs@140 sample was highly stable. In addition, the enhanced PEC water splitting potential window of the MoS₂@PVA/TNAs@140 composite for the Na₂SO₄ electrolyte is also indicated as Figure S3.

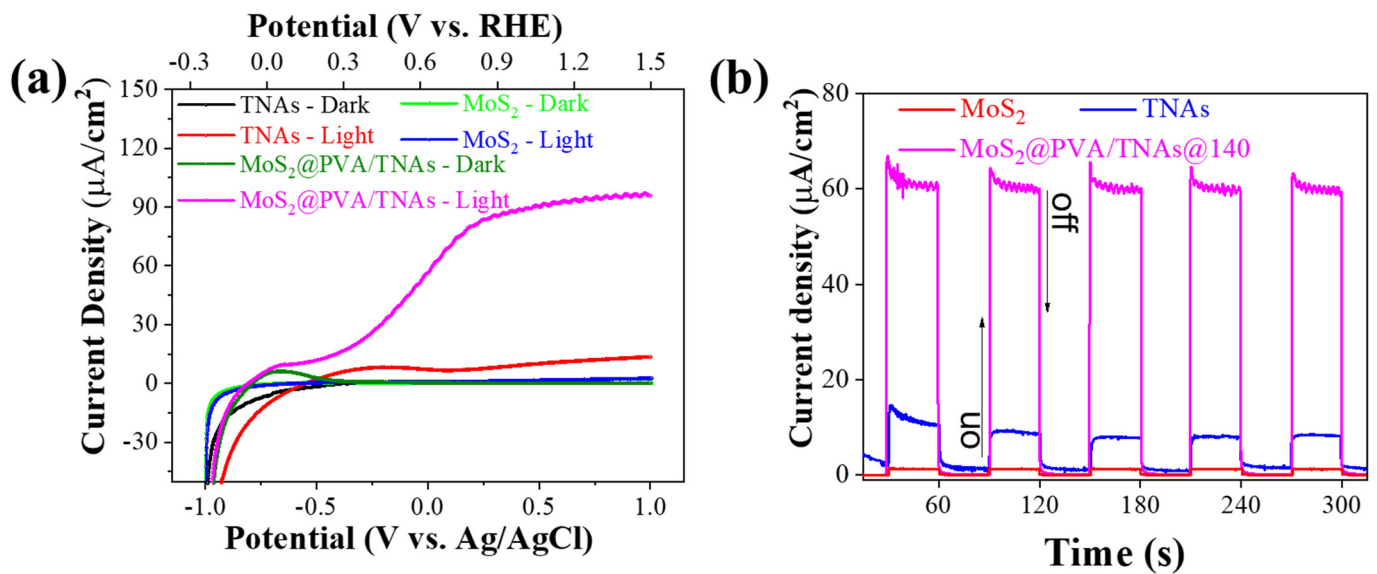


Figure 5. LSV curves of materials (a) and photoresponse ability of the materials with an applied potential at 0 V vs. Ag/AgCl (b).

The EIS plots have been obtained to study the effect of the modified material MoS₂@PVA/TNAs@140 on the charge transfer at an electrolyte/electrode interface. Figure 6a shows the single semicircle for all samples in typical Nyquist plots that express the charge transfer resistance equivalent to the polarization resistance [37]. In addition, Figure 6a shows that MoS₂ demonstrates the largest radius semicircle, followed by TNAs, indicating a weak interaction between pristine TNAs and MoS₂ with electrolytes. Furthermore, Figure 6a shows a depressed semicircle of the MoS₂@PVA/TNAs@140 sample compared to the others. This indicates a decrease in the solid-state interface layer resistance and the charge transfer resistance across the solid–liquid junction on the surface by the support of cPVA [38]. The depressed semicircle also indicates an enhancement in carrier density of MoS₂@PVA/TNAs@140 due to a decrease in resistance. This proved that the support of the π conjugation in enhancing electrical conduction increases the interaction of the electrode and the electrolyte. The Tafel slope, shown in Figure 6b, is an important parameter for evaluating the PEC activity in water splitting. The small Tafel slope value indicates a rapid increase in the reaction rate, with the applied potential corresponding to the high electrical conduction rate shown in the Nyquist plot. Figure 6b demonstrates that the MoS₂@PVA/TNAs@140 has the smallest Tafel slope value of 140.6 mV/decade, while TNAs and MoS₂ have 169.4 mV/decade and 1849 mV/decade, respectively. MoS₂@PVA/TNAs@140 shows a similar result to previous studies on MoS₂/TiO₂ materials synthesized by other methods [39]. These results indicate that the MoS₂/TNAs could be easily synthesized by combining MoS₂ with TNAs with a cPVA binding agent.

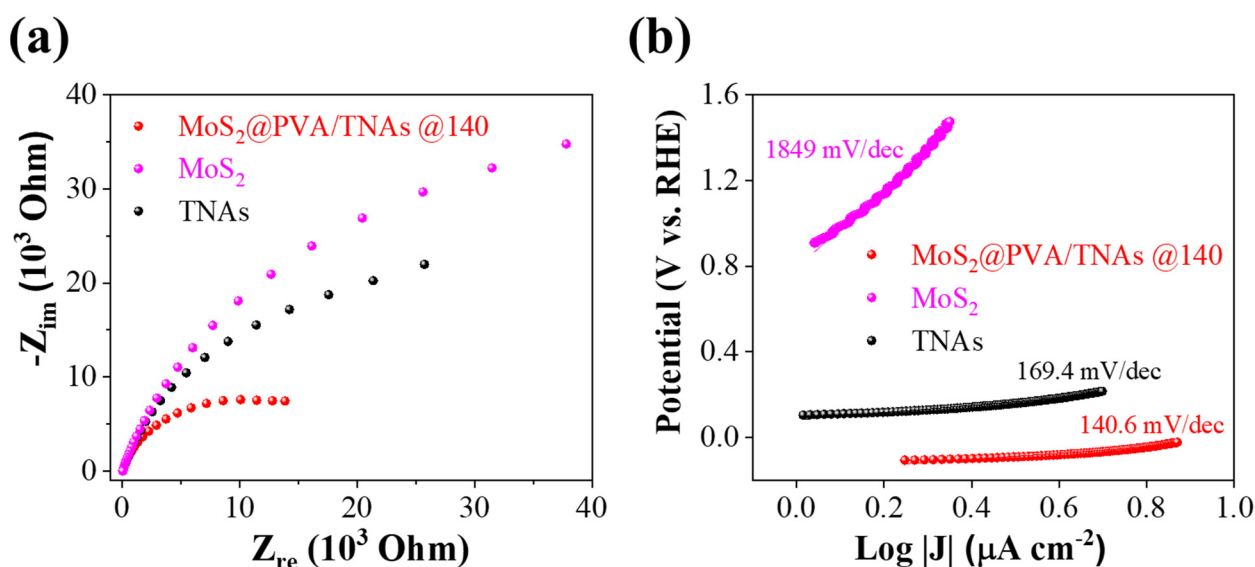


Figure 6. Electrochemical impedance spectroscopy (a) and Tafel slopes (b) of TNAs, MoS_2 , and $\text{MoS}_2@PVA/TNAs@140$.

Mott–Schottky measurements of TNAs, MoS_2 , and $\text{MoS}_2@PVA/TNAs@140$ samples have been carried out to investigate further the behavior of electrons (e^-) and holes (h^+) on the enhanced PEC activity (Figure 7a). The positive slopes in the Mott–Schottky plots indicate that all samples have n-type semiconductor properties. It is well known that the flat band potential (E_{fb}) measured from the Mott–Schottky plot for n-type semiconductors is attributed to the lowest conduction band (CB) [40,41]. The plot of $1/C^2$ vs. potential shows an intercept of the x -axis, which corresponds to the flat band potential of the material. As expected, the E_{fb} for TNAs, a well-known n-type semiconductor, is determined to be -0.89 V (V vs. Ag/AgCl, pH 7.3), which is consistent with earlier publications [42]. The E_{fb} level for the MoS_2 is determined to be -0.98 V (V vs. Ag/AgCl, pH 7.3), while the E_{fb} of $\text{MoS}_2@PVA/TNAs@140$ is found to be -0.94 V (V vs. Ag/AgCl). A negative shift of the flat band potential for $\text{MoS}_2@PVA/TNAs@140$ proves that the recombination of the $e^- - h^+$ pairs has been decreased remarkably [43]. Notably, the $\text{MoS}_2@PVA/TNAs@140$ sample shows a substantially smaller slope of the curve than the pristine TNAs and MoS_2 , indicating a significant increase in the carrier density [43,44].

Figure 7b shows the DRS spectra obtained to evaluate the energy band diagram of materials. As can be seen, the DRS spectrum of TNAs shows a strong absorption band edge at 390–395 nm regions. Meanwhile, the absorption edge of MoS_2 extends from the UV to the visible light range. The combination of TNAs and MoS_2 possesses a strong light-harvesting ability. In detail, a redshift from 395 nm to 420 nm results in an enhance in photoactivity of $\text{MoS}_2@PVA/TNAs@140$ in the visible light region. This demonstrated a crucial role of MoS_2 in increasing the photoresponse in the visible light region. Figure 7c depicts the Tauc plots of materials from which the band gaps of TNAs, MoS_2 , and $\text{MoS}_2@PVA/TNAs @140$ were calculated to be 3.2, 1.46, and 2.95 eV, respectively.

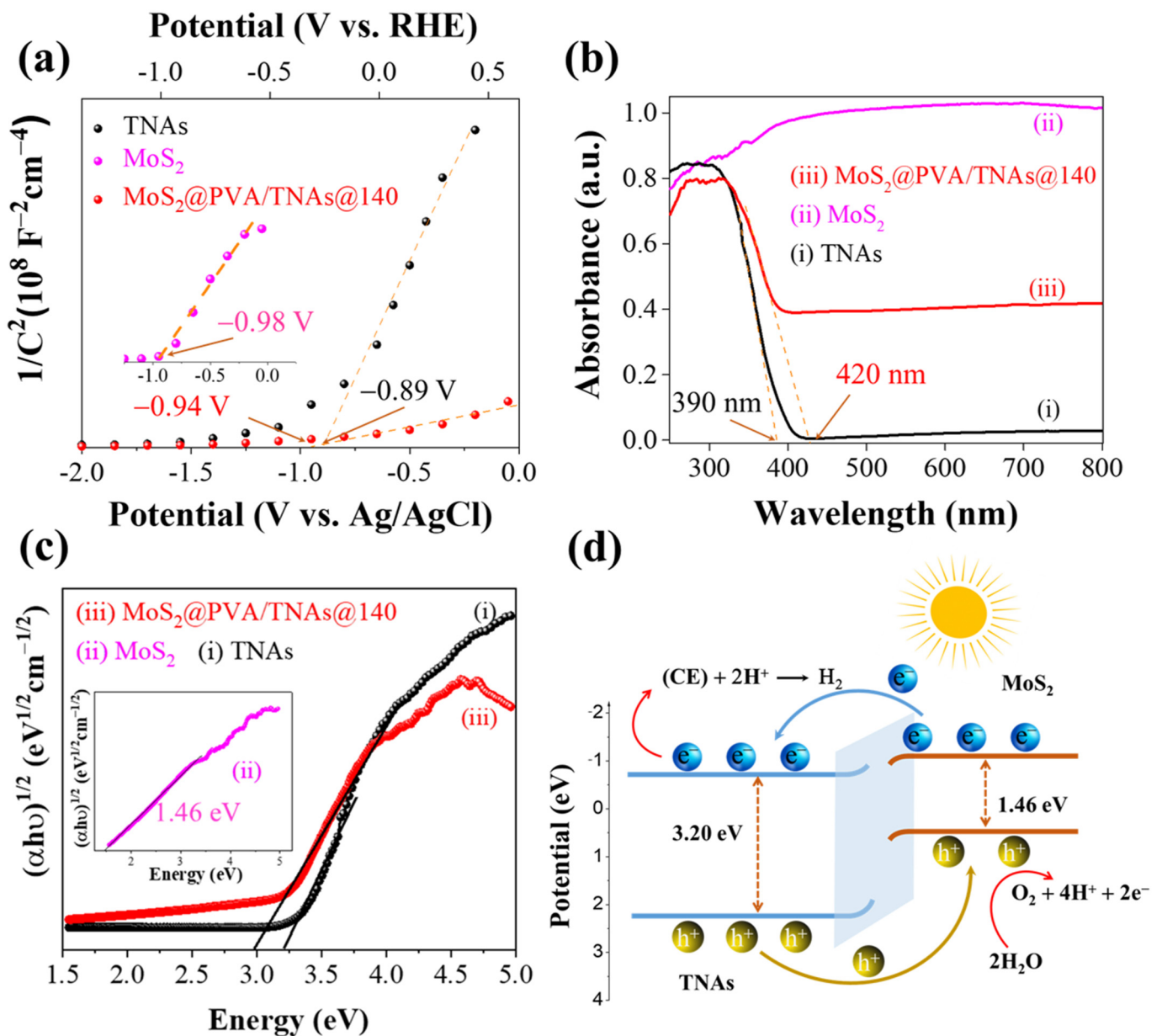
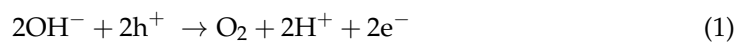


Figure 7. Mott–Schottky plots (a), DRS spectra (b), Tauc plots (c) of MoS₂, TNAs, and MoS₂@PVA/TNAs@140 and the proposed schematic illustration for the energy band alignment of the MoS₂/TNAs heterojunction (d).

According to previous studies, the bandgap of MoS₂ can be oscillated from 1.0 eV to 1.9 eV by reducing the atomic-scale thickness from bulk to monolayer [45]. Based on the DRS spectra and Mott–Schottky results, the energy band diagram of the MoS₂ and TNAs heterojunction was performed and shown in Figure 7c. The formation of this type II heterojunction could promote the migration of h⁺ and e⁻ under excitation. Additionally, this result indicated the separation of e⁻-h⁺ pairs. According to Equation (1), the h⁺ can migrate from the valence band (VB) of the TNAs to the VB of MoS₂, and then react with H₂O at the electrolyte contact.



It is commonly known that in neutral aqueous media, water splitting involves two processes. The first process is the oxidation process of water, meaning the reaction of the water with the h⁺ charges to release free H⁺ ions and O₂ [46]. At the same period, the e⁻

can move back from the CB of MoS₂ to the CB of TNAs. For the second process, free H⁺ ions are reduced at the cathode according to Equation (2):



The more H₂O oxidation reactions in Na₂SO₄ solution occur at WE, the more e⁻ and H⁺ ions are generated. This result strengthens our theory for a reason related to the enhancement of OER at the WE, which subsequently releases free H⁺ ions to react at the CE to evolve the H₂ molecule.

3. Experimental Section

3.1. Chemicals and Materials

Titanium foil (Ti, 0.25 mm of thickness, 99.6%), ammonium fluoride (NH₄F, 99.9%), ethylene glycol (C₂H₆O₂, 99.9%), sodium chloride (NaCl, 99.9%), sodium sulfate (Na₂SO₄, 99.9%), ammonium molybdate (VI) tetrahydrate ((NH₄)₆Mo₇O₂₄•4H₂O, 81–83%), N-Acetyl-L-cystine (NAC, C₅H₉NO₃S, 98.0%), thiourea (CH₄N₂S, 99.9%), polyvinyl alcohol ((C₂H₄O)_n, 99.0%), polyvinyl alcohol (PVA, 99.9%), and deionized water (Milli-Q, 18 MΩcm) were used for the preparation of the composites, solutions and/or for the rinsing steps.

3.2. Synthesis of TNAs by an Anodization Method

TNAs were synthesized by an anodization method, following the previous study [47]. First, Ti foils were cut into 1 × 2 cm and were ultrasonically washed in an acetone solution for 15 min before being soaked in an anhydrous ethanol solution for 15 min. Second, Ti foils were anodized in 2 steps at 30 V for 150 min using a DC power source in 45 mL of an electrolyte containing 40.5 mL of ethylene glycol, 4.5 mL DI water, and 0.25 mg NH₄F. Finally, the foils were washed, dried, and annealed at 400 °C for 2 h.

3.3. Synthesis of MoS₂ by a Hydrothermal Method

MoS₂ powder was synthesized by a hydrothermal method according to the previous study in ref. [48] with the following parameters: The first step was to dissolve 37.08 mg of (NH₄)₆Mo₇O₂₄•4H₂O in 40 mL of DI water in an ice bath. Following that, 65.28 mg of NAC and 30.45 mg of thiourea were added. The solution was then stirred for an hour while being purged with Ar gas to remove the oxygen gas. Subsequently, the mixture was transferred into an autoclave and kept at 200 °C for 24 h. Finally, the black precipitates were washed with ethanol and dried at 60 °C in Ar gas conditions for 12 h.

3.4. Thermal Treatment of PVA for the Formation MoS₂@PVA/TNAs Materials

Initially, 1 g PVA powder was dissolved in 99 mL DI water to form a PVA solution. Afterward, the MoS₂ powder was homogeneously dispersed in 3 mg/L PVA solutions by a sonication treatment for an hour. Finally, 0.2 mL of each solution was dropped onto the surface of TNAs and treated at different temperatures ranging from 100 °C to 180 °C in an Ar gas atmosphere. The samples were denoted as MoS₂@PVA/TNAs@100, MoS₂@PVA/TNAs@120, MoS₂@PVA/TNAs@140, MoS₂@PVA/TNAs@160, and MoS₂@PVA/TNAs@180, where the numbers represent the treatment temperatures.

To evaluate the thermal effect on the formation of cPVA structure, the pristine PVA was thermally treated from 100 to 180 °C in Ar gas, denoted as PVA@100 to PVA@180, respectively, before being analyzed by FTIR technique.

3.5. Characterizations and PEC Evaluation of Materials

FTIR spectroscopy was used to detect the vibration mode of molecules, the chemical bond of PVA, and the confirmation of the formation of cPVA using an FTIR spectrophotometer (JASCO-4700, Tokyo, Japan). X-ray diffraction (XRD) using Cu-Kα radiation (λ = 0.15405 nm, 45 kV and 40 mA) was employed to determine the crystal structures and phase state of materials. The morphology of the materials was determined by scanning elec-

tron microscopy (SEM) using FESEM equipment, Hitachi S-4800 (Tokyo, Japan) at 10.0 kV. The optical bandgap of the materials was characterized by UV-Vis diffuse reflectance spectroscopy (DRS) using a JASCO V-550 (Tokyo, Japan) spectrophotometer. The band gap of materials was determined based on an energy-dependent absorption coefficient α can be expressed by the following Equation (3):

$$(\alpha h\nu)^{\frac{1}{\gamma}} = B(h\nu - E_g) \quad (3)$$

where h , ν , E_g , and B are the Planck constant, the photon's frequency, band gap, and the constant of band gap energy. The $\gamma = 2$ for the nature of the electron transition for indirect transition band gaps of MoS₂ and TiO₂ [49–51].

The PEC characterization was conducted using a three-electrode cell with an electrochemical analyzer (BioLogic SP-200, Paris, France). In detail, the as-prepared samples were used as working electrodes (WE), a Pt wire as the counter electrode (CE), and an Ag/AgCl 3 M KCl electrode as the reference electrode (RE). An electrolyte solution of a 1 M Na₂SO₄, pH 7.3, was purged with Ar gas to remove the oxygen gas in the OER measurement. All photoactive electrodes were illuminated under a 150 W Xe lamp with a light intensity of 100 mW/cm² at the PEC cell. A cut-off filter at 380 nm of wavelength was used to avoid the UV rays. The linear sweep voltammogram (LSV) of the materials was obtained by collecting the current density according to the applied linear potential from -1 to 1 V vs. Ag/AgCl. The photoresponse ability of the materials was carried out by using a potential of 0 V vs. Ag/AgCl under the "on-off" illumination condition. The electrochemical impedance spectrum (EIS) was measured by applying a bias of the open-circuit potential over the frequency from 100 kHz to 100 mHz at the applied potential of 1 V vs. Ag/AgCl. Mott–Schottky plots were recorded from the impedance-potential test from -2 to 0.5 V vs. Ag/AgCl at 100 Hz without light irradiation. The potential V vs. Ag/AgCl was converted to the potential V vs. RHE through Equation (4) [52].

$$E_{\text{RHE}} = E_{\text{Ag/AgCl}} + E_{\text{Ag/AgCl}}^0 + 0.059 \times \text{pH} \quad (4)$$

where $E_{\text{Ag/AgCl}}^0 = 0.210$ V at 25 °C, and pH 7.3 is the pH value of the electrolyte in our experiment.

4. Conclusions

In summary, we have successfully combined MoS₂ nanoparticles and TNAs by a cPVA binding agent to enhance the PEC water splitting efficiency and the catalyst's stability. The structure of MoS₂@PVA/TNAs is well controlled by adjusting the temperature of the thermal treatment process, with the best performance was obtained at 140 °C. The results reveal that the photocurrent density of the MoS₂@PVA/TNAs@140 reaches a maximum value of 65 $\mu\text{A}/\text{cm}^2$ at 0 V vs. Ag/AgCl, which is six times higher than that of the TNAs. The combination of MoS₂ and TNAs can significantly increase the charge density, hence improving the material's conductivity. Additionally, MoS₂@PVA/TNAs@140 demonstrated excellent PEC efficiency and stability within five cycles. Moreover, by increasing conductivity, cPVA can be used as a binding agent in various semiconducting materials to enhance their properties as a stable and efficient catalyst, especially in PEC water splitting.

Supplementary Materials: The following are available online at <https://www.mdpi.com/article/10.3390/catal11070857/s1>, Figure S1: CV curve of cPVA under dark condition (Scan rate: 10 mV s⁻¹), Figure S2: LSV of pristine cPVA in the dark and under visible light irradiation., Figure S3: PEC window of MoS₂@PVA/TNAs@140.

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